

Balancing Chemical Equation Worksheet

Balance the following equations by writing in the simplest whole number coefficients in the spaces provided:

1. $__ \text{HNO}_3 + __ \text{KOH} \rightarrow __ \text{KNO}_3 + __ \text{H}_2\text{O}$
2. $__ \text{H}_2\text{SO}_4 + __ \text{LiOH} \rightarrow __ \text{Li}_2\text{SO}_4 + __ \text{H}_2\text{O}$
3. $__ \text{HF} + __ \text{NaOH} \rightarrow __ \text{NaF} + __ \text{H}_2\text{O}$
4. $__ \text{H}_3\text{PO}_4 + __ \text{Ca}(\text{OH})_2 \rightarrow __ \text{Ca}_3(\text{PO}_4)_2 + __ \text{H}_2\text{O}$
5. $__ \text{HBr} + __ \text{NH}_4\text{OH} \rightarrow __ \text{NH}_4\text{Br} + __ \text{H}_2\text{O}$

Write in the formulas for the products in the following double replacement reactions and then balance them:

1. $__ \text{Na}_2\text{CO}_3 + __ \text{HCl} \rightarrow __ \text{NaCl} + __ \text{H}_2\text{O} + __ \text{CO}_2$
2. $__ \text{Mg}(\text{OH})_2 + __ \text{H}_2\text{C}_2\text{O}_4 \rightarrow __ \text{MgC}_2\text{O}_4 + __ \text{H}_2\text{O}$
3. $__ \text{K}_2\text{S} + __ \text{H}_2\text{SO}_4 \rightarrow __ \text{K}_2\text{SO}_4 + __ \text{H}_2\text{S}$
4. $__ \text{Ca}(\text{OH})_2 + __ \text{H}_3\text{PO}_4 \rightarrow __ \text{Ca}_3(\text{PO}_4)_2 + __ \text{H}_2\text{O}$
5. $__ \text{NH}_4\text{OH} + __ \text{HNO}_3 \rightarrow __ \text{NH}_4\text{NO}_3 + __ \text{H}_2\text{O}$

