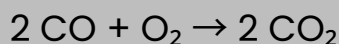


# Balancing Chemical Equations Worksheet

Answer the following questions about the chemical equation shown below:



- What are the reactants?
- What is the product?
- What do we call the number "2" in front of the CO and CO<sub>2</sub>?
- Why is there no coefficient for O<sub>2</sub>?
- How many carbon atoms are needed to produce two CO<sub>2</sub> molecules?
- How many oxygen molecules are needed to produce two CO<sub>2</sub> molecules?
- How many oxygen atoms are needed to produce two CO<sub>2</sub> molecules?
- Write the word equation that you would use to describe this reaction.
- Use words in a sentence, not formulas or an arrow, to describe the reaction.

Balance the following chemical equations:

- $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- $\text{NH}_3 + \text{O}_2 \rightarrow \text{NO} + \text{H}_2\text{O}$
- $\text{CH}_4 + \text{Cl}_2 \rightarrow \text{CH}_3\text{Cl} + \text{HCl}$
- $\text{P}_4 + \text{Cl}_2 \rightarrow \text{PCl}_3$

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