



# Balancing Chemical Equations Worksheet

## Decomposition or Analysis

1. \_\_\_ HgO  $\rightarrow$  \_\_\_ Hg + \_\_\_ O<sub>2</sub>
2. \_\_\_ NaHCO<sub>3</sub>  $\rightarrow$  \_\_\_ Na<sub>2</sub>CO<sub>3</sub> + \_\_\_ CO<sub>2</sub> + \_\_\_ H<sub>2</sub>O
3. \_\_\_ CaCO<sub>3</sub>  $\rightarrow$  \_\_\_ CaO + \_\_\_ CO<sub>2</sub>
4. \_\_\_ NH<sub>4</sub>NO<sub>3</sub>  $\rightarrow$  \_\_\_ N<sub>2</sub>O + \_\_\_ H<sub>2</sub>O
5. \_\_\_ KClO<sub>4</sub>  $\rightarrow$  \_\_\_ KCl + \_\_\_ O<sub>2</sub>

## Single Replacement

1. \_\_\_ Ag + \_\_\_ H<sub>2</sub>S  $\rightarrow$  \_\_\_ Ag<sub>2</sub>S + \_\_\_ H<sub>2</sub>
2. \_\_\_ Cu + \_\_\_ AgNO<sub>3</sub>  $\rightarrow$  \_\_\_ Cu(NO<sub>3</sub>)<sub>2</sub> + \_\_\_ Ag
3. \_\_\_ Zn + \_\_\_ HCl  $\rightarrow$  \_\_\_ ZnCl<sub>2</sub> + \_\_\_ H<sub>2</sub>
4. \_\_\_ Fe + \_\_\_ CuSO<sub>4</sub>  $\rightarrow$  \_\_\_ FeSO<sub>4</sub> + \_\_\_ Cu
5. \_\_\_ Cl<sub>2</sub> + \_\_\_ KBr  $\rightarrow$  \_\_\_ KCl + \_\_\_ Br<sub>2</sub>