



## Balancing Chemical Equations Worksheet

1.  $\_\_ \text{N}_2 + \_\_ \text{H}_2 \rightarrow \_\_ \text{NH}_3$  (Ammonia Synthesis)
2.  $\_\_ \text{C}_2\text{H}_4 + \_\_ \text{O}_2 \rightarrow \_\_ \text{CO}_2 + \_\_ \text{H}_2\text{O}$   
(Combustion of Ethylene)
3.  $\_\_ \text{Al} + \_\_ \text{Fe}_2\text{O}_3 \rightarrow \_\_ \text{Al}_2\text{O}_3 + \_\_ \text{Fe}$  (Thermite Reaction)
4.  $\_\_ \text{Pb}(\text{NO}_3)_2 \rightarrow \_\_ \text{PbO} + \_\_ \text{NO}_2 + \_\_ \text{O}_2$   
(Decomposition of Lead(II) Nitrate)
5.  $\_\_ \text{CuSO}_4 + \_\_ \text{BaCl}_2 \rightarrow \_\_ \text{BaSO}_4 + \_\_ \text{CuCl}_2$   
(Double Displacement Reaction)
6.  $\_\_ \text{H}_2\text{SO}_4 + \_\_ \text{NaOH} \rightarrow \_\_ \text{Na}_2\text{SO}_4 + \_\_ \text{H}_2\text{O}$   
(Neutralization Reaction)
7.  $\_\_ \text{KClO}_3 \rightarrow \_\_ \text{KCl} + \_\_ \text{O}_2$  (Decomposition of Potassium Chlorate)
8.  $\_\_ \text{CH}_4 + \_\_ \text{Cl}_2 \rightarrow \_\_ \text{CH}_3\text{Cl} + \_\_ \text{HCl}$   
(Halogenation of Methane)