

# Balancing Chemical Equations Worksheet

## Section 1: Basic Reactions

1. \_\_\_  $\text{H}_2$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{H}_2\text{O}$
2. \_\_\_  $\text{Na}$  + \_\_\_  $\text{Cl}_2$   $\rightarrow$  \_\_\_  $\text{NaCl}$
3. \_\_\_  $\text{N}_2$  + \_\_\_  $\text{H}_2$   $\rightarrow$  \_\_\_  $\text{NH}_3$
4. \_\_\_  $\text{C}$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{CO}_2$
5. \_\_\_  $\text{Mg}$  + \_\_\_  $\text{HCl}$   $\rightarrow$  \_\_\_  $\text{MgCl}_2$  + \_\_\_  $\text{H}_2$

## Section 2: Intermediate Reactions

1. \_\_\_  $\text{Fe}$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{Fe}_2\text{O}_3$
2. \_\_\_  $\text{CaCO}_3$   $\rightarrow$  \_\_\_  $\text{CaO}$  + \_\_\_  $\text{CO}_2$
3. \_\_\_  $\text{Al}$  + \_\_\_  $\text{Br}_2$   $\rightarrow$  \_\_\_  $\text{AlBr}_3$
4. \_\_\_  $\text{NH}_3$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{NO}$  + \_\_\_  $\text{H}_2\text{O}$
5. \_\_\_  $\text{CH}_4$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{CO}_2$  + \_\_\_  $\text{H}_2\text{O}$

## Section 3: Advanced Reactions

1. \_\_\_  $\text{Pb}(\text{NO}_3)_2$  + \_\_\_  $\text{KI}$   $\rightarrow$  \_\_\_  $\text{PbI}_2$  + \_\_\_  $\text{KNO}_3$
2. \_\_\_  $\text{C}_6\text{H}_{12}\text{O}_6$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{CO}_2$  + \_\_\_  $\text{H}_2\text{O}$
3. \_\_\_  $\text{Zn}$  + \_\_\_  $\text{CuSO}_4$   $\rightarrow$  \_\_\_  $\text{ZnSO}_4$  + \_\_\_  $\text{Cu}$