

Counting Atoms Worksheet

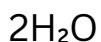
Name: _____ Grade: _____

Instructions:

Look at each chemical formula below and:

- Identify each element.
- Count the number of atoms for each element.
- Pay attention to subscripts and coefficients (if included).

Example:



$$\text{H} = 4$$

$$\text{O} = 2$$

1. H_2O
 $\text{H} = \underline{\hspace{1cm}}$
 $\text{O} = \underline{\hspace{1cm}}$

6. 3NH_3
 $\text{N} = \underline{\hspace{1cm}}$
 $\text{H} = \underline{\hspace{1cm}}$

2. CO_2
 $\text{C} = \underline{\hspace{1cm}}$
 $\text{O} = \underline{\hspace{1cm}}$

7. $\text{Al}_2(\text{SO}_4)_3$
 $\text{Al} = \underline{\hspace{1cm}}$
 $\text{S} = \underline{\hspace{1cm}}$
 $\text{O} = \underline{\hspace{1cm}}$

3. NaCl
 $\text{Na} = \underline{\hspace{1cm}}$
 $\text{Cl} = \underline{\hspace{1cm}}$

8. 4KNO_3
 $\text{K} = \underline{\hspace{1cm}}$
 $\text{N} = \underline{\hspace{1cm}}$
 $\text{O} = \underline{\hspace{1cm}}$

4. $\text{C}_6\text{H}_{12}\text{O}_6$
 $\text{C} = \underline{\hspace{1cm}}$
 $\text{H} = \underline{\hspace{1cm}}$
 $\text{O} = \underline{\hspace{1cm}}$

9. $\text{Mg}_3(\text{PO}_4)_2$
 $\text{Mg} = \underline{\hspace{1cm}}$
 $\text{P} = \underline{\hspace{1cm}}$
 $\text{O} = \underline{\hspace{1cm}}$

5. $\text{Ca}(\text{OH})_2$
 $\text{Ca} = \underline{\hspace{1cm}}$
 $\text{O} = \underline{\hspace{1cm}}$
 $\text{H} = \underline{\hspace{1cm}}$

10. $\text{Ba}(\text{NO}_3)_2$
 $\text{Ba} = \underline{\hspace{1cm}}$
 $\text{N} = \underline{\hspace{1cm}}$
 $\text{O} = \underline{\hspace{1cm}}$

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Answer Key

